

Stoichiometry Study Guide

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The relative masses were obtained by multiplying the atomic ratios and atomic masses. You can see that a sample of N_2O weighing 44.02 grams contains 28.02 g of nitrogen and 16.00 g of oxygen. The mass percent of each element is calculated from its relative mass divided by the sum of the relative masses. Chemical compounds with integral atomic ratios, like nitrous oxide, are described as stoichiometric compounds, and they permit many simple calculations.

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Page 1/16. Download Ebook Study Guide Chemistry Stoichiometry Answer Key. Study Guide Chemistry Stoichiometry Answer The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. _____ 2. Stoichiometry is based on the law of conservation of mass.

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Whip out your tool belt, and grab the instrument labeled stoichiometry. It is time to get quantitative. It doesn't matter if the units given for reactants (or products) are in moles, grams, liters, or any other units imaginable. The key to solving stoichiometry problems is to first convert any given unit to moles. Repeat: always convert to moles.

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The best definition for stoichiometry is the simple one: it's a way to figure out how much stuff you're going to make in a chemical reaction, or how much stuff you'll need to make a chemical reaction do what you want. When we put it that way, stoichiometry isn't so bad. We can deal with the crazy name if it's that simple.

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