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with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) → 2NH<sub>3</sub>(g) ... mc06se cFMsr i-vi - Chapter 9 Stoichiometry Section 1 Answers  
The balancing numbers are known as the coefficients of the reaction.

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through "File info" CorrectionKey=NL-A

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the line to the left. Show all your work in the space provided. 1. 88%  
The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) ...

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relationships of elements in compounds Reaction Stoichiometry -  
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chemical reaction I. Page 4/15

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