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When H forms a bond with H O to form the hydronium ion H O , this bond is called a coordinate covalent bond because

- a. both bonding electrons come from the oxygen atom.
- b. it forms an especially strong bond.
- c. the electrons are equally shared.
- d. the oxygen no longer has eight

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Key: 109.5 degrees that results when a central atom forms four bonds directed toward the center of a regular tetrahedron. the mixing of several atomic orbitals to form the same total number of equivalent hybrid orbitals.

Chapter 8: Covalent Bonding Crossword

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Covalent Bonding Covalent Bonding

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Covalent Bond pages 240 – 247 Practice

Problems page 244 Draw the Lewis

structure for each molecule. 1. PH₃ H

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Key: H—H H P respectively, for single,
double, and triple P — — 2. H 2 S H H H
— H S S ...

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Choose the best answer and write its letter on the line. 1. A bond in which each atom contributes two electrons is a. a double covalent bond. b. an ionic bond. c. a polar covalent bond. d. a coordinate covalent bond. 2. The electron dot structure for hydrogen sulfide, H_2S ...

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sharing of electrons occurs the attachment
between atoms is called answercovalent
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Covalent Bonding - Chapter 8. 1.

Covalent Bonding Or How I Learned to
Love Sharing (But Remember, File
Sharing is Illegal) 2. As you should
remember, ionic compounds are solids at
room temperatures that have one ion strip

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Key: the electron (s) from the other elements ' electron cloud.

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Chapter 8 Covalent Bonding and
Molecular Structure 8-11. nuclei. This
results in stronger attractive forces

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Key: Between electrons and nuclei, decreasing the distance between the nuclei. A carbon-carbon single bond has a bond order of 1 and is longer than a carbon-carbon double bond with a bond order of 2.

Chapter 8: Covalent Bonding and
Molecular Structure

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Chapter 8 Covalent Bonding Section 81
Molecular Compounds ...

Chapter 8: Covalent Bonding. Matter takes many forms in nature: In this chapter, we are going to learn to distinguish the type of compound that we have already studied, the “ ionic compound ” (which contains oppositely-

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charged particles: metal cations and non-metal anions), from a different type of compound – a “ molecular compound ” .

Chapter 8: Covalent Bonding

Chapter 8 - Covalent Bonding/Polarity

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CHOICE 1. ANS: D PTS: 1 DIF: L3

REF: p. 248 | p. 249 OBJ: 8.4.1 Describe how electronegativity values determine the charge distribution in a polar molecule.

MSC: application 2. ANS: B PTS: 1 DIF: L1 REF: p. 251

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