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base neutralization pogil answers acid base neutralization pogil answers Acid-Base Neutralization Reactions Exercises 1. Based upon the information presented in the key of the Model, draw reactants and products that form when an H^+ ion is added to an OH^- ion. 2.

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Answers

Acid-Base Neutralization Reactions Exercises 1. Based upon the information presented in the key of the Model, draw reactants and products that form when an H^+ ion is added to an OH^- ion. 2. What would happen if solution A and solution B were mixed? Explain your answer. 3. Classify the solution that forms in Exercise 2 as acidic, basic, or ...

Acid – Base Neutralization Reactions -
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In the reaction below identify which of the reactants is an acid and which is a base: +

$\text{H}_3\text{O}^+(\text{aq}) + \text{OH}^-(\text{aq})$ 2. Consider the atomic

structure of the H^+ ion. Complete the table

below indicating the correct number of each

subatomic particle. Composition of the H^+

ion (H^+) Subatomic Number of particle

subatomic particles Protons Electrons

Neutrons 3.

Mrs. Zuberbuehler - Mrs. Zuberbuehler

13. All acid—base reactions have two

conjugate acid—base pairs. One conjugate

acid—base pair in the reaction in Model 3 is

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H O'/H O. List the other acid—base pair in the reaction. 14. Why is HCO considered the "acid" part of the pair in the reaction in Model 3? 15. is CO considered the "base" part of the pair in the reaction in Model 3? 16.

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The sodium bicarbonate will have a pH greater than 7 (between 8 and 14) because it is a base.

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